NAME \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ GRADED: TRY THIS: ION REVIEW

SCORE: 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1 0

DIRECTIONS: This is graded. Yes, you can use your notes (pages 36 – 43 should be of help). Be effective in your communication. If we can’t read it, it is wrong… So consider using scrap paper for Time Period 1…

This is being broken down into TWO TIME PERIODS.

Time Period 1) **Independent Work**: You and you alone are to **work with your notes** to answer as many

questions as you can. You are not to speak with anyone else. Talking (and listening) loses

points. You will not be warned, just docked. (10 minutes)

Time Period 2) Group Work: You and ONE partner, can work together to attack any question you or s/he could

not get … or those about which you were feeling fragile. You are welcome to change answers,

in this group time. Ms. M or Mr. D will assign groups and/or create a group of 3 if needed.

(5 minutes)

1) Tell me what you know about the term ION. Consider providing a definition … or bullet ideas that show

your understanding of the term. You do not need to use complete sentences, but be clear. (5 pts)









(you can use the back if needed)

2) When we use the term ***ELECTROLYTE***, it is referring to **an ion, under certain circumstance and/or with**

**a special ability**. When we speak of electrolytes … what do we mean and/or what can they do? We had

a demonstration on this … and we did use Gatorade…thanks o Dillon! (3 pts) (you can use the back if

needed)

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Questions 3-5) Use your periodic table &/or notes. Complete the table. (6 pts)

|  |  |
| --- | --- |
| Atom | Ion |
| 82Pb0  p= , e- = | Pb4**+**  p= e- = 1 |
| 35Br0  p= , e- = | Br**-**1  p= , e- = |
| 6C0  p= , e- = | C4**-**  p= , e- = |

3

4

5

Turn this over and go on….

For questions 6 - 8 select the choice which most correctly completes each statement. Put the number of your choice in the space next to each question. A choice may used once, more than once or not at all.

1) negative in overall charge 2) neutral in overall charge 3) positive in overall charge

place answer here

\_\_\_6) When an atom loses electrons (negative charge), the resulting ion should be \_\_\_\_\_\_\_\_\_\_\_\_(1 pt)

\_\_\_7) When an atom gains electrons (negative charge), the resulting ion should be \_\_\_\_\_\_\_\_\_\_\_\_\_(1 pt)

\_\_\_8) When a +1 cation gains an electron, the resulting species should be \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(1 pt)

**Use this space if you need more for question 1:**

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**Use this space if you need more for question 2:**

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